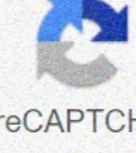


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**SUBMIT**

## Chapter No. 1

## BASIC CONCEPTS

## TEXT BOOK EXERCISE

- Q1. Select the most suitable answer from the given ones in each question.
- (i) The mass of one mole of electrons is  
 (a) Properties which depend upon mass  
 (b) Arrangement of electrons in orbital  
 (c) Chemical properties  
 (d) The extent to which they may be affected in electromagnetic field
- (ii) Which of the following statements is not true?  
 (a) isotopes with even atomic masses are comparatively abundant  
 (b) isotopes with odd atomic masses and even atomic number are comparatively abundant  
 (c) atomic masses are average masses of isotopes.  
 (d) Atomic masses are average masses of isotopes proportional to their relative abundance
- (iii) Many elements have fractional atomic masses, this is because  
 (a) The mass of the atom is itself fractional  
 (b) Atomic masses are average masses of isotopes  
 (c) Atomic masses are average masses of isotopes.  
 (d) Atomic masses are average masses of isotopes proportional to their relative abundance
- (iv) The mass of one mole of electrons is  
 (a) 0.08mg (b) 0.55mg (c) 0.184mg (d) 1.673mg
- (v) 27g of Al will react completely with how much mass of O<sub>2</sub> to produce Al<sub>2</sub>O<sub>3</sub>  
 (a) 8g of oxygen (b) 16g of oxygen  
 (c) 32g of oxygen (d) 24g of oxygen
- (vi) The number of moles of CO<sub>2</sub> which contain 8.0 g of oxygen.  
 (a) 0.25 (b) 0.50 (c) 1.0 (d) 1.50
- (vii) The largest number of molecules are present in  
 (a) 3.6g of H<sub>2</sub>O (b) 4.8g of C<sub>2</sub>H<sub>5</sub>OH  
 (c) 2.8 g of CO (d) 5.4g of N<sub>2</sub>O<sub>5</sub>
- (viii) One mole of SO<sub>2</sub> contains  
 (a) 6.02x10<sup>23</sup> atoms of oxygen

Only Chemistry Discussion

GCSE CHEMISTRY  
ATOMS ELEMENTS AND  
COMPOUNDS MIXTURES

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BASIS	COMPOUNDS	MIXTURE
Composition	Fixed	Variable
Separation	Components cannot be separated by physical method	Can be separated by physical methods
Properties	Compound has different property than its constituents	All the components retain their properties
Chemical Bond	Components are chemically bonded	Not chemically bonded
Chemical Reaction	Involved chemical reaction in formation	No Chemical reaction
Melting and boiling points	have sharp and fixed MP and BP Water H <sub>2</sub> O	do not have sharp and fixed MP or BP Salt and water
Examples	Methane CH <sub>4</sub> Hydrogen Chloride HCl	Sugar and Water Oil and Water

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## Chapter 20- Chemical Bonds

- I. Stability of Bonding
- Some properties combine chemically and no longer have the same *properties* they did before forming a compound.
  - A(n) **chemical formula** is composed of symbols and subscripts indicating the number of atoms of an element in a compound.
  - Atoms form compounds when the compound is more *stable* than the separate atoms.
    - Noble gases are more **chemically stable** than other elements because they have a complete outer energy level.
    - Atoms can lose, gain, or *share* electrons to get a stable outer energy level.
    - A(n) **chemical bond** is the force that holds atoms together in a compound.
- II. Types of Bonds
- A(n) **ion** is charged particle because it has more or fewer electrons than protons.
    - When an atom *loses* an electron, it becomes a positively charged ion; a superscript indicates the charge.
    - When an atom *gains* an electron, it becomes a negatively charged ion.
  - An ionic compound is held together by the **ionic bond**, the force of attraction between opposite charges of the ions.
    - The result of this bond is a **neutral** compound.
    - The sum of the charges on the ions is **zero**.
  - Molecules** are neutral particles formed as a result of sharing electrons.
    - A **covalent bond** is the force of attraction between atoms sharing electrons.
    - Atoms can form double or triple **bonds** depending on whether they share two or three pairs of electrons.
    - Electrons shared in a molecule are held **more closely** to the atoms with the larger nucleus.
    - A(n) **polar molecule** has one end that is slightly negative and one end that is slightly positive although the overall molecule is neutral.
    - In a(n) **nonpolar molecule** electrons are shared equally.
- III. Writing Formulas and Naming Compounds
- Chemists use **symbols** from the periodic table to write formulas for compounds.
  - Binary Compounds**- composed of two elements
    - Oxidation number**- how many electrons an atoms has gained, lost, or shared to become stable
    - Use oxidation numbers and their least common multiples to write **formulas**.
      - When writing formulas, remember that the compound is **neutral**.
      - A formula must have the correct number of positive and negative ions so the charges **balance**.
    - Use the name of the first element, the root name of the second element, and the suffix *-ide* to write the **name** of the binary ionic compound.
  - Polyatomic ion**- positively or negatively charged, covalently bonded groups of atoms.
    - The compound contains **three** or more elements.
    - To write names, write the name of the **positive** ion first; then write the name of the **negative** ion.
    - To write **formulas**, use the oxidation numbers, their least common multiple, and put parentheses around the polyatomic ion before adding a subscript.
  - Hydrate**- compound with water chemically attached to its ions
  - Name binary covalent compounds by using **prefixes** to indicate how many atoms of each element are in the compound.







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